

shape and the ends introduced into the glass tubes. These are then gently heated until the wax melts. The ends of the steel wire should reach down as far as possible into the glass tubes to insure strength.

GRASSELLI, N. J., Feb. 18th, 1907.

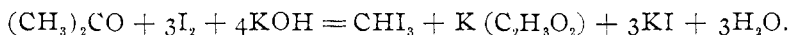
IVAR HOEL.

A Rapid Method for the Estimation of Acetone.—In order to be able to estimate rapidly the acetone content of the distillate from crude wood naphtha during rectification, the author has adapted the well-known Messinger method (*Ber.*, 21, 3368; see also Kebler, *This Journal*, 19, 316 (1897),) so that the burette readings give the number of grams of acetone in 100 cc. (sometimes expressed as per cent.) directly.

Procedure.—Put 10 c.c. of 2N sodium hydroxide solution into a wide-mouth glass-stoppered bottle, then add exactly 1 c.c. of the alcohol to be examined, shaking; next add 50 c.c. of the standard iodine solution while shaking; after three minutes acidify with dilute sulphuric acid (use about 2N H_2SO_4); now add starch indicator and titrate back with standard thiosulphate solution. The number of cubic centimeters of iodine solution used by the acetone gives the number of grams of the latter in 100 c. c.

For the back titration the writer uses a burette graduated from 0 to 50 c. c. in $\frac{1}{10}$, beginning at the bottom; after titration, the burette reading gives grams of acetone in 100 c.c. directly, avoiding a subtraction of the reading from 50.

In accordance with the equation:



0.01g of acetone requires 0.1312 g. of iodine. A standard iodine solution is, accordingly, prepared to contain 131.2 g. of iodine in one liter; the standard thiosulphate solution is prepared of equivalent strength.

This method is intended for spirits in which the acetone is high, but does not exceed 50%. In cases where the acetone does not run over 25%, a 25 c.c. burette, graduated from the bottom, should be used for the thiosulphate solution, and of course only 25 c.c. of the iodine solution are used. For solution containing 5% or less of acetone, 10 c.c. of the sample may be taken, using the 50 c.c. burette; readings in this case are 0.1% for each c.c. iodine used.

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The test for formaldehyde in milk by Leach's modification of the hydrochloric acid and ferric chloride test.—Recently the government inspector took some samples of milk from our house and had them examined by the government pure food chemist. He reported that he found formaldehyde in the milk. We tested some of the milk from the same source (but not the same lot) and found no formaldehyde. Later he reported